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Solution: Carbon in graphite form has the electron configuration of $1s^2 2s^2 2p^2$ which only allows for 3 bonds. Additions of pressure and heat, hybridization occurs to the point that the electron moves from the 2s to the 2p orbital ($1s^2 2s^1 2p^3$) allowing access for 4 bonds to occur.

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Solution: Let a be the length of the sides of the tetrahedral cube and r be the radius of the two types of atoms. $a^3 = r^3 + r^3$ $a = 2r$ $C H a = 2rC + 2rH = 2(0.77 \times 10^{-8} \text{ cm} + 0.46 \times 10^{-8} \text{ cm}) = 1.42 \times 10^{-8} \text{ cm}$ $-8 -8 3 3 7. 2-20$
The compound aluminum phosphide (AlP) is a compound semiconductor having mixed ionic and covalent bonding.

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